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Relation	



Single-nanometer sized low-valence metal hydroxide crystals: synthesis via epoxide-mediated alkalization and assembly towards functional mesoporous materials

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Metal oxides of low-valence metal cations (M(II) and M(III); M = Al, Cr, Mn, Fe, Co, Ni, Cu, etc.) are essential in catalysis, sustainable energy, information technology and environmental applications. In addition, the corresponding metal hydroxides present a 2D crystalline structure of hydroxide layers and interlayer anions, and have attracted attention due to their properties, such as reversible and fast intercalation, deintercalation, and anisotropic transfer of ions.¹ The wide variety of combination of the metal cations composing the 2D hydroxide layers and the interlayer anions leads to enhanced electrochemical,² magnetic,³ optical⁴ and catalytic⁵ properties. These phases in turn can be transformed into other layered phases such as sulfides that present interesting electronic properties.⁶

In order to fully exploit the surface and bulk properties dictated by the nano/mesostructure of these low-valence phases, it is desirable to process them as materials with highly accessible surface, controlled morphology and tailored crystallite size. The possibility of shaping mesoporous metal oxides and hydroxides with high specific surface area, controlled pore size and tunable surface functionality is particularly interesting in this framework.⁷

Although considerable progress has been made in the controlled synthesis of mesoporous silica and non-silica oxides, a limitation exists in the access to mesoporous low-valence metal oxides, hydroxides and derived phases. To date, only very few synthetic paths are reported, and the mesostructures obtained are seldom robust.⁸ The difficulty of obtaining these phases can be traced back to the condensation chemistry of M(II)/M(III) centers, which present very fast crystallization and aggregation, and therefore hinder the co-assembly with supramolecular templates. These strategies rely mostly on the use of hard templates, or on obtaining textural porosity derived from interparticle spaces.⁹ An alternative way to obtain highly ordered mesoporous phases is the use of pre-formed fine nanoparticles with high controllability in single-nm scale.¹⁰ Nanoparticles presenting a given crystalline phase and well-defined shape can co-assemble with the supramolecular tem-

plates as a nano-building block (NBB), avoiding the lengthy steps associated to the hard template approach. However, the synthesis of single-nm sized crystals of layered metal hydroxides and their use as NBBs remain a considerable challenge. In general, a low degree of supersaturation is favored for synthesizing hydroxides to avoid aggregation, yielding coarse crystals whose size is too large to be used as NBB. Quite recently, we reported the synthesis of nanometric Ni(II)-Al(III)-type layered double hydroxide (LDH).¹¹ The reaction successfully yielded 5 nm LDH which could be used as NBBs of mesoporous LDHs, whereas Al(III) was found to be a mandatory element therein to achieve the developed reaction scheme involving a gelation-deflocculation transition. It still remains a challenge to achieve a chemically versatile synthesis of hydroxide NBBs made of desired cationic species.

Herein, we demonstrate a simple synthesis approach toward single-nm sized crystals that work as NBB to form a great variety of mesoporous hydroxides and derived phases. The nanocrystals with diverse chemical compositions (layered low-valence metal hydroxides, α -M(OH)₂ (M = Mn(II), Fe(II), Co(II), Ni(II), and Cu(II)), and other (hydr)oxides such as Al(OH)₃, CrO(OH), β -FeO(OH), ZrO₂, and SnO₂) were prepared in an aqueous ethanol solution. The nanocrystal formation was achieved by inducing a high supersaturation via the epoxide-mediated alkalization in a concentrated metal salt solution. A variety of carboxylic acids were used as growth controlling agents, permitting to finely control the size of low-valence metal hydroxide nanocrystals and to attain high stability against aggregation. Mesoporous hydroxide materials with nanocrystalline walls were obtained by co-assembly of crystalline NBBs and F127 templates through the evaporation-induced self-assembly (EISA) process as films, powders and aerosols. Functional crystalline mesostructures with desired shape dictating chemical/physical properties have been thus successfully obtained. We discuss mesoporous α -Ni(OH)₂ films as an example, in which the advantages of nanocrystalline nature and mesoporous structure were demonstrated by

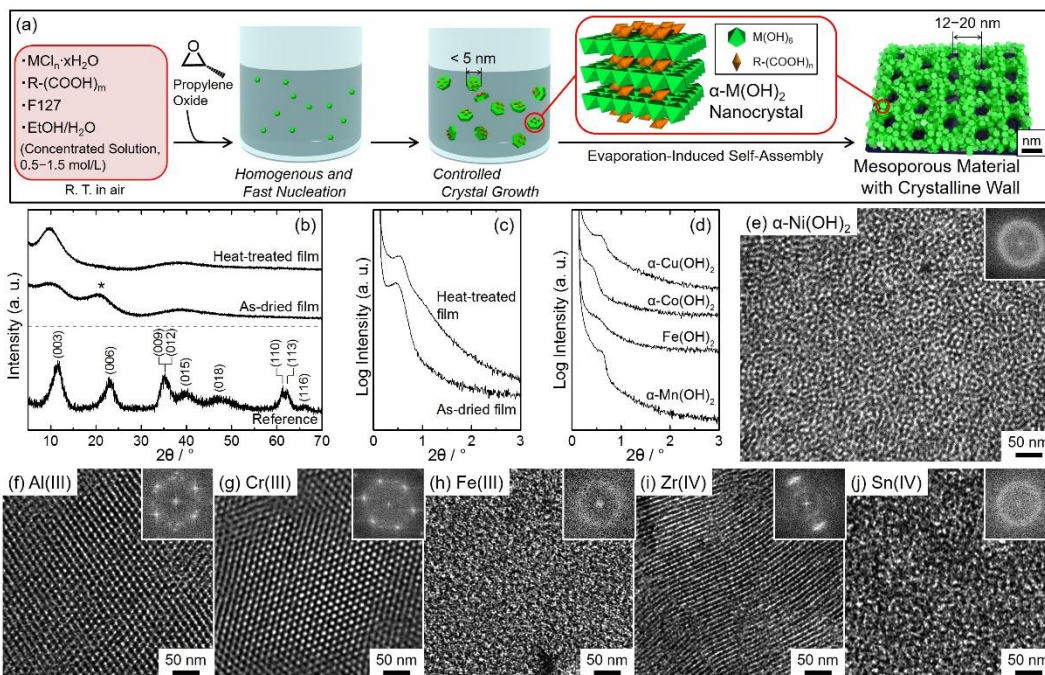


Figure 1. (a) Schematic illustration of our approach to prepare single-nm sized nanocrystals toward mesoporous materials with crystalline walls. (b) PXRD and (c) SAXS patterns of as-dried and heat-treated ($250\text{ }^\circ\text{C}$, in air) films prepared from $\alpha\text{-Ni(OH)}_2$ nanocrystals intercalated with $\text{C}_2\text{H}_4(\text{COOH})_2$. The reference pattern in (b) corresponds to $\text{C}_2\text{H}_4(\text{COOH})_2$ intercalated Ni-Al-type LDH prepared via a standard coprecipitation method. (d) SAXS patterns of as-dried films prepared from $\alpha\text{-M(OH)}_2$ ($\text{M} = \text{Mn(II)}, \text{Fe(II)}, \text{Co(II)}, \text{and Cu(II)}$) nanocrystals intercalated with $\text{C}_3\text{H}_6(\text{COOH})_2$. TEM images of heat-treated films prepared from (e) $\alpha\text{-Ni(OH)}_2$ intercalated with $\text{C}_2\text{H}_4(\text{COOH})_2$, and (f)–(j) various (hydr)oxides nanocrystals. Heat treatment temperatures (in air): (e) $250\text{ }^\circ\text{C}$, (f)–(j) $350\text{ }^\circ\text{C}$. Insets of (e)–(j) are FFT patterns generated from the respective images. * Overlapped diffraction peaks derived from $\alpha\text{-Ni(OH)}_2$ and an impurity interstratified Ni(OH)_2 .

pesudomorphic transformations to functional phases, and the improvement of pseudocapacitor performance.

Figure 1(a) shows a schematic illustration of the fabrication of these mesoporous films with crystalline walls. The starting mixture is an aqueous ethanol solution containing metal chloride hydrate (its concentration is 5–100 times higher than those in general methods to achieve high degree of supersaturation¹¹), carboxylic acid and F127. Adding propylene oxide to the Ni-containing starting mixture results in an homogenous and fast pH increase through the epoxide ring-opening reaction with the nucleophilic species.¹² Along with the pH increase, ionic Ni species were consumed by the nucleation and growth of $\alpha\text{-Ni(OH)}_2$ nanocrystals; these processes are evidenced by the decrease of electric conductivity (Figure S1(a)) and X-ray diffraction (XRD) of reacting solution (Figure S1(b)). Small angle X-ray scattering (SAXS) measurements confirmed the formation of crystals with a gyration radius of 2.6 nm (Figure S1(c)) which is 10-times smaller than without carboxylic acid (26.1 nm). These nanocrystals can be further used as NBB for functional mesostructures. Upon coating, surfactant (F127) self-assembled into ordered mesostructures¹³ together with NBB by EISA process. Mesoporous thin films with hydroxide nanocrystalline walls were obtained after removal of F127.

Figure 1(b) shows powder XRD (PXRD) patterns of the as-dried and the heat-treated films of $\alpha\text{-Ni(OH)}_2$ with $\text{C}_2\text{H}_4(\text{COOH})_2$ as a growth controlling agent. A broad peak around $2\theta = 10^\circ$ corresponds to a d -spacing of 8.6 Å, which is coincident with (001) interplanar distance of Ni-Al layered double hydroxide prepared by a standard coprecipitation technique. The broad peaks indicate that the obtained film is con-

stituted of hydroxide nanocrystals. The crystallite size along c axis, out-of-plane direction of hydroxide layers, calculated from Scherrer's equation is less than 3 nm. The peak ascribed to (110) plane, in-plane direction of hydroxide layers, was not clearly detected due to the small lateral size of crystals. The film retained the (001) diffraction ($d_{001} = 8.8\text{ \AA}$) even after the removal of F127 by heat treatment at $250\text{ }^\circ\text{C}$ (the removal of F127 was confirmed by FT-IR, see Figure S2), indicating that the obtained mesostructured $\alpha\text{-Ni(OH)}_2$ film was thermally stable. Figure 1(c) shows SAXS patterns of the as-dried and the heat-treated films. Periodic lengths calculated from the diffraction peaks are 17.9 nm and 16.1 nm for the as-dried and the heat-treated films, respectively. Transmission electron microscopy (TEM) and scanning electron microscopy (SEM) confirmed a templated mesoporous structure with a well-defined pore size (Figure 1(e) and S3). A N_2 adsorption/desorption isotherm of the mesoporous $\alpha\text{-Ni(OH)}_2$ film is classified into a typical type IV shape with a hysteresis loop of H1 (Figure S4), which is characteristic of a cylinder type mesoporous structure. The mesoporous film has a large BET specific surface area ($141\text{ m}^2\text{g}^{-1}$), a large pore size (9.2 nm) and a large pore volume ($0.30\text{ cm}^3\text{g}^{-1}$) compared to the film prepared without F127 (hereafter denoted as a nonporous film). The wall thickness of the mesostructure was calculated as 6.9 nm from the result of SAXS and N_2 sorption. The size of the $\alpha\text{-Ni(OH)}_2$ NBB is small enough compared to the wall thickness, which is the key for the synthesis of the templated mesoporous structure.

The present NBB approach is applicable to various chemical systems including low-valence metal ions; Mn(II), Fe(II), Co(II), Cu(II), Al(III), Cr(III), Fe(III), Zr(IV), and Sn(IV),

yielding nanocrystals of α -Mn(OH)₂, α -Co(OH)₂, Fe(OH)₂, α -Cu(OH)₂, Al(OH)₃, CrO(OH), β -FeO(OH), ZrO₂, and SnO₂, respectively (Figure S5(a)). The SAXS patterns and TEM images confirm that the films have ordered mesostructures with periodicities ranging from 14.4 to 20.7 nm (Figure 1(d)–(j), S5(b) and Table S1). Our approach toward mesoporous films can also be coupled with a spray drying technique. Micron-sized spherical particles with defined mesopores were successfully obtained by the aerosol-assisted EISA process (Figure S6). Spray drying is a high yield and scalable processing methodology desired for industrial applications.¹⁴

The compositional versatility of both hydroxide and anion layers offers a considerable advantage in applications, that is tunability of chemical and physical characteristics of mesoporous layered metal hydroxide materials. For instance, the composition of hydroxide layers and interlayers is reported to affect capacitance of layered low-valence metal hydroxides.^{2,15} Figure 2(a) shows PXRD patterns of mesoporous films prepared from α -Ni(OH)₂ NBBs intercalated with different dicarboxylic acids (C_nH_{2n}(COOH)₂, n = 1–5). With increasing the alkyl chain length, the (001) peak shifts to lower angles. The increase of *d*₀₀₁ is 0.58 Å per one chain unit. Even when (001) *d*-spacing was modulated by using different carboxylic acids, mesoporous structures were successfully formed (Figure S7(a)–(e)). Various mono-, di- and tri-carboxylic acids can be intercalated in α -M(OH)₂ nanocrystals which can be assembled into mesoporous structures as NBB (Table S1). The tunability of interlayer anions as well as cationic species is a significant advance over previous reports on nanometric metal hydroxides.^{11,17}

As Stucky and coworkers reported,¹⁶ acetic acid can control the inorganic condensation of M(IV) (M = Ti, Zr, Si and mixtures) so that inorganic species could lead to co-assembly with supramolecular template. In the present study, we found that carboxylic acids used as a growth controlling agent played multiple roles. As is evidenced by infrared spectra and XRD patterns (Figure S2 and S8(a)), the carboxylic acids coordinate and adsorb on the surface/edges of 2D hydroxide layers, which inhibits

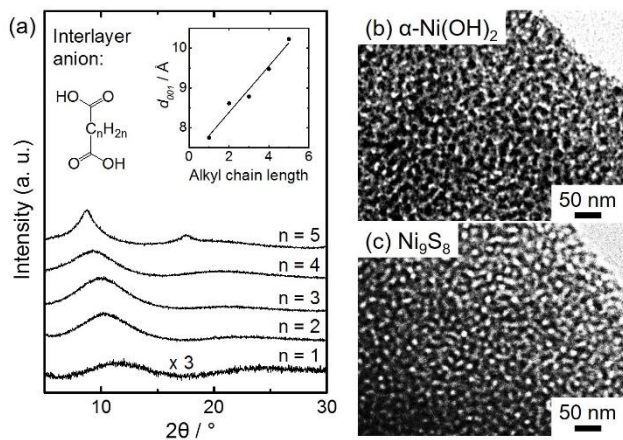


Figure 2. (a) PXRD patterns of films prepared from α -Ni(OH)₂ NBBs intercalated with C_nH_{2n}(COOH)₂ (n = 1–5) after the heat treatment at 250 °C in air. Inset shows *d*₀₀₁ vs alkyl chain length (as carbon number). TEM images of mesoporous films prepared from α -Ni(OH)₂ NBB intercalated with C₂H₃SH(COOH)₂ (b) before and (c) after phase transformation to Ni₉S₈.

extensive crystal growth of hydroxide in the in-plane (110) direction.¹⁷ This allows to form nanocrystals small enough for NBB approach. The carboxylic acids are intercalated in the interlayer cavity as well as the crystal surface. The intercalated anionic species then can be used as chemical sources for pseudomorphic transformation as discussed later. In addition, surface carboxylic acid groups help to stabilize nanocrystals against aggregation (Figure S8(b) and (c)). The carboxylic acid is therefore indispensable to control the structural feature, as well as chemical and surface features of the α -M(OH)₂ NBB toward the formation of mesoporous structure.

Since layered hydroxides can be easily transformed to metals, oxides, and metal organic frameworks through pseudomorphic transformation,^{1b,18} the simultaneous control over the nano/mesostructure and the chemical composition of layered hydroxide offers a versatile pathway toward tuning the physical and chemical properties of these materials. The obtained α -Ni(OH)₂ films with a variety of interlayer anions were subjected to further heat treatments (Table S2). The interlayer anions work as chemical sources to tune the composition of transformed phase such as NiO, Ni, Ni/Ni₃C, Ni/C and Ni₉S₈ (Figure S9). By tuning the heat treatment condition, phase transformation was achieved without collapse of mesopore structure (Figure 2(b) and (c)). The present nanocrystals can serve as precursors for useful and functional mesoporous materials composed of crystalline walls such as MnO₂, NaCoO₄, Li[Mn_xCo_yNi_z]O₄.¹⁹

As a proof-of-concept study, an electrochemical property was investigated for α -Ni(OH)₂ films. Cyclic voltammetry (CV) and galvanostatic charge-discharge measurements were performed on the mesoporous and the nonporous α -Ni(OH)₂ films. The CV curves of both α -Ni(OH)₂ films show a pair of redox peaks (Figure 3(a), S10(a) and (b)), which is a typical feature of Ni(OH)₂ pseudocapacitor with Faradic reaction of Ni(OH)₂ + OH⁻ ↔ NiOOH + H₂O + e⁻.²⁰ The non-linear galvanostatic

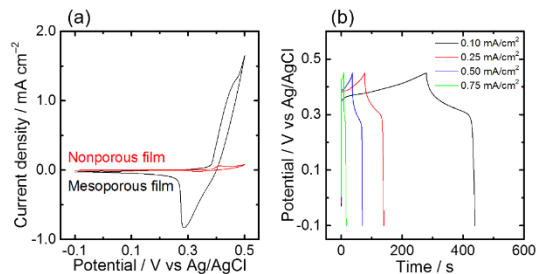


Figure 3. (a) CV curves of mesoporous and nonporous α -Ni(OH)₂ films at a scan rate of 10 mV/s. (b) Galvanostatic charge-discharge curves of the mesoporous α -Ni(OH)₂ film at various current densities. Intercalated carboxylic acid: C₂H₄(COOH)₂.

charge-discharge curves also confirm the pseudocapacitance behavior due to quasi-reversible redox reaction (Figure 3(b) and S10(c)). The area-specific capacitances calculated from the discharge curves of 0.10 mA/cm² were 29.7 and 0.176 mF/cm² for mesoporous and nonporous films, respectively. A difference of the specific surface areas of these films is only 4.4 times, which cannot account for the significant difference in the capacitances. The considerable increase of capacitance (169 times) is due to improved electrolyte mobility and high accessibility to the active sites by the introduction of mesopore of adequate size (3–10 nm).²¹ The mesoporous film showed a high mass-specific capacitance (926 F/g), specific energy

(31.4 Wh/kg) and specific power (3466 W/kg) at discharge current of 14.0 A/g (Table S3). As well as mesoporosity, nanocrystalline nature contributes to enhancement of pseudocapacitor performance. Fabricating α -Ni(OH)₂ electrode with nanocrystals allows to introduce a high density of grain boundaries which is known to work as an efficient diffusion pass of electrolyte ions.²² The nanocrystalline mesoporous α -Ni(OH)₂ films reported here is a promising material for a high performance capacitor.

In summary, single-nm sized α -Ni(OH)₂ crystals in an aqueous ethanol solution were synthesized through a simple one batch process using an epoxide-mediated alkalization with carboxylic acids. Co-assembly of crystalline NBBs and supramolecular template by EISA lead to ordered mesoporous thin films, xerogels, and aerosol-derived powders. This NBB assembly approach can be extended to other transition low- and high-valence metal ions (Mn(II), Fe(II), Co(II), Cu(II), Al(III), Cr(III), Fe(III), Zr(IV), and Sn(IV)) which were so far difficult to obtain, leading to nanocrystalline mesoporous thin films with a variety of chemical compositions. The compositional versatility enables the phase conversion from α -Ni(OH)₂ to oxides, metals, metal/carbide or metal/carbon composites, and sulfides. The introduction of ordered mesopores enhanced the pseudocapacitance performance of α -Ni(OH)₂ film. The NBB approach reported here offers a promising route to obtain transition metal-based mesoporous materials/composites which are useful in a wide variety of research fields, such as catalysis, adsorbents, thermo/chemical electric or magnetic devices, and sensing.

ASSOCIATED CONTENT

Supporting Information. Experimental details, XRD patterns, FT-IR spectra, N₂ sorption, SEM and TEM images, Raman spectrum, electrochemical analysis.

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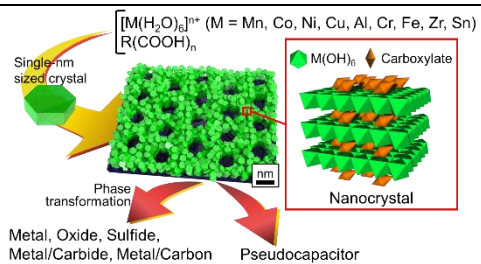
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Single-nanometer sized crystals of low-valence metal hydroxides were synthesized by epoxide-mediated alkalization incorporated with carboxylic acid. Co-assembly of nanocrystals and supramolecular templates led to form mesostructured thin films, xerogels, aerosol powders. Layered hydroxides transformed to functional phases. Mesoporous α -Ni(OH)₂ films showed increase of pseudocapacitance by the introduction of mesostructures.