

広島大学学術情報リポジトリ
Hiroshima University Institutional Repository

| | |
|------------|---|
| Title | Ultraviolet Photodissociation Spectroscopy of Cold K ⁺ Calix[4]arene Complex in the Gas Phase |
| Author(s) | Inokuchi, Yoshiya; Soga, Kazuki; Hirai, Kenta; Kida, Motoki; Morishima, Fumiya; Ebata, Takayuki |
| Citation | Journal of Physical Chemistry A , 119 (31) : 8512 - 8518 |
| Issue Date | 2015-08-06 |
| DOI | 10.1021/acs.jpca.5b05328 |
| Self DOI | |
| URL | https://ir.lib.hiroshima-u.ac.jp/00046471 |
| Right | Copyright (c) 2015 American Chemical Society This document is the Accepted Manuscript version of a Published Work that appeared in final form in Journal of Physical Chemistry A, copyright c American Chemical Society after peer review and technical editing by the publisher. To access the final edited and published work see https://doi.org/10.1021/acs.jpca.5b05328 . |
| Relation | |



Ultraviolet Photodissociation Spectroscopy of Cold K^+ •Calix[4]arene Complex in the Gas Phase

**Yoshiya Inokuchi,* Kazuki Soga, Kenta Hirai, Motoki Kida, Fumiya Morishima,
and Takayuki Ebata**

*Department of Chemistry, Graduate School of Science, Hiroshima University,
Higashi-Hiroshima, Hiroshima 739-8526, Japan*

E-mail: y-inokuchi@hiroshima-u.ac.jp

Abstract

The cooling of ionic species in the gas phase greatly simplifies the UV spectrum, which is of special importance to study the electronic and geometric structures of large systems, such as bio-related molecules and host-guest complexes. Many efforts have been devoted to achieving the ion cooling with a cold quadrupole Paul ion trap (QIT), but one problem was insufficient cooling of ions (up to ~ 30 K) in the QIT. In this study, we construct a mass spectrometer for ultraviolet photodissociation (UVPD) spectroscopy of gas-phase cold ions. The instrument consists of an electrospray ion source, a QIT cooled with a He cryostat, and a time-of-flight mass spectrometer. Giving a great care for the cooling condition, we can achieve ~ 10 K for the vibrational temperature of ions in the QIT, which is estimated from UVPD spectra of the benzo-18-crown-6 (B18C6) complex with potassium ion, K^+ •B18C6. Using this setup, we measure a UVPD spectrum of cold calix[4]arene (C4A) complex with potassium ion, K^+ •C4A. The spectrum shows a very weak band

and a strong one at 36018 and 36156 cm^{-1} , respectively, accompanied by many sharp vibronic bands in the 36000–36600 cm^{-1} region. In the geometry optimization of the $\text{K}^+\bullet\text{C4A}$ complex, we obtain three stable isomers: one endo and two exo forms. On the basis of the total energy and UV spectral patterns predicted by density functional theory calculations, we attribute the structure of the $\text{K}^+\bullet\text{C4A}$ complex to the endo isomer (C_2 symmetry), in which the K^+ ion is located inside the cup of C4A. The vibronic bands of $\text{K}^+\bullet\text{C4A}$ at 36018 and 36156 cm^{-1} are assigned to the $\text{S}_1(\text{A})\text{--}\text{S}_0(\text{A})$ and $\text{S}_2(\text{B})\text{--}\text{S}_0(\text{A})$ transitions of the endo isomer, respectively.

Keywords: calix[4]arene, alkali metal, ion trap, electrospray, ultraviolet, conformation

*To whom correspondence should be addressed.

1. Introduction

Calixarenes (CAs) are macrocycle compounds consisting of phenol units, and exhibit encapsulation and self-assembly.¹ Their cup-like forms provide cavities to include guest species and show the guest selectivity, which can be controlled by the ring size and introduction of functional groups. Alkali metal ion-CA complexes were extensively studied as simple systems of charged guest-CA complexes. Izatt and coworkers examined cation transport through a liquid membrane system containing CA carriers and discussed the selectivity of alkali metal ions by CAs.² In condensed phase, the conformation of CA complexes with alkali metal ions was determined by X-ray diffraction and NMR spectroscopy.³ The stoichiometry of CA complexes and the guest selectivity of CAs in solutions were studied by mass spectrometric methods, coupled with electrospray ionization (ESI) and matrix-assisted laser desorption/ionization (MALDI).⁴⁻⁸ In organic chemistry, Shinkai and coworkers synthesized a number of CA derivatives and succeeded to control their conformations, such as cones and partial cones.⁹⁻¹³ They also demonstrated that the metal ion selectivity of CAs is highly dependent on their conformations. Furthermore, Haino and coworkers synthesized polymeric compounds formed with CAs and guest species such as fullerenes.¹⁴ Laser-based spectroscopy for complexes in the gas phase, coupled with cooling techniques such as free jet expansion and cold ion traps, is also useful for determining the complex structure without the interference from solvent. The cooling of neutral or ionic species in the gas phase greatly simplifies the UV spectrum, which makes it possible to separate features from different stable conformers and examine the electronic and geometric structures individually. Concerning spectroscopy of CAs in the gas phase, we have investigated the structure of cold inclusion complexes of calix[4]arene (C4A) with neutral guests such as rare gas atoms,

water, ammonia, molecular nitrogen, methane, and acetylene, under free jet conditions.¹⁵⁻¹⁸ The experimental and theoretical results suggest that all the inclusion complexes of C4A preferentially form endo complexes (the guests are inside the C4A cup). In addition to the ability of inclusion, the electronic structure of C4A is also of great interest. Bare C4A has C_4 symmetry with its four phenol moieties, which interact with each other via through space interaction. As a result, C4A exhibits weak S_1-S_0 transition at 35357 cm^{-1} and strong S_2-S_0 transition at 35521 cm^{-1} .¹⁵ It will be quite interesting how the inclusion of ionic species affects the geometric and electronic structure of C4A. For electronic spectroscopy of ionic species, ultraviolet photodissociation (UVPD) spectroscopy is a very powerful technique. Boyarkin, Rizzo, and their coworkers reported pioneering works for UVPD spectroscopy of cold protonated amino acids with a cold, 22-pole ion trap.¹⁹⁻²⁰ Kim and coworkers extensively studied UVPD spectroscopy of ion encapsulation complexes of benzo-crown ethers.²¹⁻²⁶ Recently, we have investigated benzo-crown ether complexes with metal ions and their hydrated species, coupled with IR-UV double resonance spectroscopy, to determine the structure and the number of conformers and to shed light on the origin of the ion selectivity.²⁷⁻³¹

In the present study, we construct a mass spectrometer for UVPD spectroscopy, equipped with an ESI ion source and a cold, quadrupole Paul ion trap (QIT). Firstly we determine the ion temperature in the QIT by measuring UVPD spectra of benzo-18-crown-6 (B18C6) complex with potassium ion, $K^+\bullet B18C6$. In our previous paper, we have measured a UVPD spectrum of cold $K^+\bullet B18C6$ in a cold, 22-pole ion trap.²⁸ We estimate the ion temperature of $K^+\bullet B18C6$ in the QIT by comparing the UVPD spectra using the QIT and the 22-pole ion trap. Then we observe a UVPD spectrum of C4A complex with potassium ion, $K^+\bullet C4A$. We

determine the geometric and electronic structure of the $K^+ \cdot C4A$ complex from the UVPD spectrum, with the aid of quantum chemical calculations.

2. Experimental and computational methods

Figure 1 shows a schematic drawing of a mass spectrometer used in study. Ion complexes are produced continuously at atmospheric pressure by an ESI source. A stainless needle (Hamilton 21033A) is connected to a syringe (Hamilton 1010TLL) through a 1/8 in. Teflon tube, and a solution is carried from the syringe to the needle by a syringe pump (KD Scientific 780-100E) with a flow rate of ~ 0.2 mL/h. A high DC voltage (~ 3 kV) is applied to the needle for the electrospray. A vaporization tube (a stainless tube with an outer diameter of 1/16 in. and a length of ~ 300 mm) is situated at the entrance of the vacuum chamber. The tube is held by two copper blocks, surrounded by a rubber heater, and heated to ~ 100 °C. After passing through the tube and a skimmer with a diameter of 1 mm, ions are introduced into the first octopole ion guide (OPIG). At the exit of the OPIG, a gate electrode is situated to bunch the continuous ion current. This electrode is kept grounded for ~ 100 ms, then a pulsed potential of -30 V with a width of ~ 200 μ s is applied for the ions to exit the first OPIG. The pulsed ion beam is guided by the second OPIG and introduced into a QIT (Jordan TOF Products C-1251). The QIT is bolted in a copper box, which is connected to the second stage of a He cryostat (Sumitomo Heavy Industries RDK-408D2) and cooled to ~ 4 K. The QIT is used in an open configuration; the ring electrode and the end caps are held with four threaded rods and ceramic spacers, and ceramic rings between the ring electrode and the end caps are not used. The copper box is surrounded by a copper shield, which is moderately cooled by the first stage of the He cryostat. A He

gas line is attached to the first stage of the cryostat to precool He buffer gas, then the buffer gas is continuously introduced into the QIT. The ions are stored in the QIT for ~ 95 ms and cooled translationally and internally by the collision with the cold He buffer gas. Ions other than parent ions of interest are removed from the QIT by an RF potential applied to the entrance end cap, as was done by Kang et al.³² The parent ions are then irradiated by a UV laser, inducing the dissociation of the ions. After ~ 1 μ s of the UV excitation, resulting fragment ions are accelerated to a home-made time-of-flight mass spectrometer by pulsed potentials applied to the exit end cap and a potential switch located at the exit of the trap, and detected by a multichannel plate (MCP).³³ An output from the MCP is fed into a digital storage oscilloscope (LeCroy Wave Runner 604Zi) or a multi-channel scaler/average (Stanford Research Systems SR430). A signal averaged by the oscilloscope or counted by the scaler/averager is transferred to a personal computer through a USB or a GPIB interface. Yields of the fragment ions are normalized by the intensity of the UV laser, and UVPD spectra of the parent ions are obtained by plotting the normalized yields of the fragment ions against wavenumber of the UV laser. In the experiments of $\text{K}^+\bullet\text{B18C6}$ and $\text{K}^+\bullet\text{C4A}$, we use a solution containing potassium chloride and B18C6 or C4A (~ 100 μM each) dissolved in methanol.

In UVPD experiments described above, the RF voltages to the OPIGs are generated by RF power supplies, which are constructed on the basis of the RF circuit by Jones and Anderson.³⁴ The QIT is driven by a power supply for QITs (Jordan TOF Products D-1203). The pulsed voltage to the exit end cap for ejecting the fragment ions is also supplied by the same power supply. A function generator (Stanford Research Systems DS345) is used to generate the RF voltage to the entrance end cap for removing unnecessary ions from the QIT. The pulsed voltages to the gate electrode

and the potential switch are generated by high-voltage pulse generators (DEI PVX-4140). All the pulsed components are controlled by a digital delay/pulse generator (Berkeley Nucleonics Corporation Model 575). The temperature of the copper box holding the QIT is measured by a silicon diode (Lake Shore Cryotronics DT-670B-CU) and a cryogenic temperature controller (Lake Shore Cryotronics Model 325). To ensure good thermal contact in the cold components, we use cryogenic high vacuum grease (M&I Materials Apiezon N) on all the contact surfaces of the copper components and the ceramic spacers of the QIT.

The tunable UV light is obtained by second harmonic generation (SHG) of a fundamental output of a pulsed dye laser (Continuum ND6000) pumped by the third harmonics of a Nd:YAG laser (Continuum Surelite II), with a repetition rate of 10 Hz. For the SHG, we use a commercial servo system (Inrad Autotracker II) with a KDP crystal. The UV laser is focused into the QIT loosely by using a quartz lens with a focal length of 3000 mm. A typical output energy of the UV laser used in this study is ~ 0.2 mJ/pulse.

We also perform quantum chemical calculations for the $K^+ \cdot C4A$ complex and bare C4A. The geometry optimization is performed at the M05-2X/6-31+G(d) level of theory using the GAUSSIAN09 program package.³⁵ The transition energy and the oscillator strength are obtained theoretically by time-dependent density functional theory (TD-DFT) calculations at the M05-2X/6-31+G(d) level.

3. Results and discussion

First we estimate the ion temperature in our cold QIT by observing UV spectra of $\text{K}^+\bullet\text{B18C6}$. Figure 2 shows the UVPD spectra of the $\text{K}^+\bullet\text{B18C6}$ complex in the $35900\text{--}36300\text{ cm}^{-1}$ region; these spectra are measured by detecting the fragment K^+ ion. In our previous study, we have measured the UVPD spectrum of $\text{K}^+\bullet\text{B18C6}$ using a cold, 22-pole ion trap (Fig. 2a).²⁸ The origin band appears at 36101 cm^{-1} , and a weak hot band is seen at 36071 cm^{-1} . We estimated the ion temperature in Fig. 2a to be $\sim 10\text{ K}$ from the relative intensity and frequency of the hot band.²⁸ The UVPD spectra in Figs. 2b and 2c are measured in the present study. When the ion cooling is not so sufficient (Fig. 2b), the relative intensity of the hot band at 36071 cm^{-1} to the origin band is higher than that in Fig. 2a; the vibrational temperature of $\text{K}^+\bullet\text{B18C6}$ in Fig. 2b is estimated as $\sim 35\text{ K}$, which is a typical value with cold Paul ion traps.^{21, 36-37} In addition, the width of each vibronic band in Fig. 2b is larger than that of the cold spectrum in Fig. 2a. Broad features on the lower frequency side of each band maximum were ascribed to hot bands accompanied by the hot band at 36071 cm^{-1} .²¹ However, by taking a great care of avoiding the heat flow from electric connections and the He gas line and by optimizing experimental conditions such as the trapping time and the intensity of the UV laser, we can finally obtain a colder UV spectrum, as shown in Fig. 2c. The width of each band becomes sharper than that in Fig. 2b, and no noticeable hot band is seen at 36071 cm^{-1} . From the comparison of the spectra in Figs. 2a and 2c, we estimate the ion temperature in our QIT (Fig. 2c) to be $\sim 10\text{ K}$. This result indicates that ions can be cooled to $\sim 10\text{ K}$ even by using Paul ion traps. One can find the difference in the relative intensity of the progression above 36140 cm^{-1} to the origin band between Figs. 2a and c. This is probably due to the saturation effect on the UVPD spectrum in Fig. 2a. In the measurement of the spectrum in Fig. 2a, the UV laser power was ~ 1.5

mJ/pulse (though the photon density in the 22-pole ion trap was not measured), which overly enhanced the relative intensity of the weak bands above 36140 cm^{-1} to the origin band. For the UVPD spectrum in Fig. 2c, we decrease the UV power as low as possible to avoid the saturation (the intensity of the fragment K^+ ion in Fig. 2c is about $\sim 20\%$ of that in Fig. 2b). However, since the intensity of the vibronic band at 36126 cm^{-1} relative to the origin band at 36101 cm^{-1} , which was used for the evaluation of the vibrational temperature, does not change so much between Figs. 2a and 2c, the previous temperature estimation of the $\text{K}^+\cdot\text{B18C6}$ complex in Fig. 2a as $\sim 10\text{ K}$ is still reliable.²⁸

Then we observe a UVPD spectrum of the $\text{K}^+\cdot\text{C4A}$ complex by using our cold QIT. Figure 3 shows the UVPD spectrum of the $\text{K}^+\cdot\text{C4A}$ complex in the $35700\text{--}36600\text{ cm}^{-1}$ region. The spectrum is measured by monitoring the yield of the fragment K^+ ion. Since the ion intensity of the $\text{K}^+\cdot\text{C4A}$ complex is much weaker than that of the $\text{K}^+\cdot\text{B18C6}$ complex in our experiment, the signal to noise ratio of the UVPD spectrum (Fig. 3) is not good compared to that of $\text{K}^+\cdot\text{B18C6}$ (Fig. 2). In spite of the disadvantage of the ion intensity, the cooling of the ions to $\sim 10\text{ K}$ enables us to observe weak but sharp vibronic bands clearly for $\text{K}^+\cdot\text{C4A}$. A very weak but reproducible band is observed at 36018 cm^{-1} with a bandwidth of $\sim 1.8\text{ cm}^{-1}$ (FWHM). Since no band is observed on the lower frequency side, the band at 36018 cm^{-1} can be assigned to the origin band of the $\text{S}_1\text{--}\text{S}_0$ transition. Jet-cooled C4A shows the origin band of the $\text{S}_1\text{--}\text{S}_0$ transition at 35357 cm^{-1} ; the attachment of K^+ ion shifts the electronic transition of C4A to the blue by $\sim 660\text{ cm}^{-1}$.¹⁵ Such a large blue shift is observed also for the $\text{K}^+\cdot(\text{dibenzo-18-crown-6})$ complex.²⁷ As seen in Fig. 3, low-frequency progressions are observed from the origin and other low-frequency bands with an interval of $\sim 38\text{ cm}^{-1}$. The spectrum is quite congested in the $36000\text{--}36400\text{ cm}^{-1}$ region, but a strong

band clearly appears at 36156 cm^{-1} , highlighted by an arrow in Fig. 3. In addition, another strong band is found at 36515 cm^{-1} .

Based on calculated results of stable isomers and electronic transitions, we attribute the structure of the $\text{K}^+\bullet\text{C4A}$ complex in our experiment to an endo isomer. Figure 4 displays the stable structure of C4A and $\text{K}^+\bullet\text{C4A}$ calculated at the M05-2X/6-31+G(d) level of theory. Bare C4A has a C_4 structure with all the phenol components identical (C4A-I in Fig. 4a).¹⁵ In the case of the $\text{K}^+\bullet\text{C4A}$ complex, the most stable structure (KC4A-I in Fig. 4b) has its K^+ ion inside the cup (an endo form). The K^+ ion in the second and the third most stable structures (KC4A-II and KC4A-III) is located outside of the C4A cup (exo forms), attached to one of the benzene rings and to two of the four oxygen atoms, respectively. The total energy of these isomers relative to that of the most stable one is quite large (97 and 101 kJ/mol), suggesting that the $\text{K}^+\bullet\text{C4A}$ complex has KC4A-I form. This structural assignment of $\text{K}^+\bullet\text{C4A}$ is confirmed by the results of the TD-DFT calculations. Table 1 shows the calculated transition energies and oscillator strengths for C4A and $\text{K}^+\bullet\text{C4A}$. We have demonstrated in our previous study that the relative transition energy of stable isomers obtained by TD-DFT calculations can be used for the structural determination, whereas the absolute values are not consistent to the observed energies.²⁸ As seen in Table 1, isomers KC4A-I and KC4A-III show the blue shift of the $\text{S}_1\text{-S}_0$ transition compared to that of bare C4A (C4A-I). In the case of KC4A-III, the oscillator strengths to the S_1 , S_2 , and S_3 states have the same order of magnitude. For KC4A-I, in contrast, the oscillator strength to the S_1 state is substantially weaker (almost zero) than that to the S_2 and S_3 states, which reproduces the spectral features of the $\text{K}^+\bullet\text{C4A}$ complex in Fig. 3 much better than KC4A-III. Hence we conclude that the $\text{K}^+\bullet\text{C4A}$ complex in our experiment has KC4A-I form. The binding energy between K^+ and C4A for KC4A-I

is calculated to be 1.95 eV (15700 cm⁻¹); this is consistent to the experimental result that the fragment K⁺ ion is formed by the UV excitation of K⁺•C4A in the 36000–36600 cm⁻¹ region.

In KC4A-I (Fig. 4b), two of the four benzene rings hold the K⁺ ion cooperatively, changing the symmetry from C₄ (bare C4A) to C₂. In Fig. 4, the tilting angle of the benzene rings with the principal axis (C₄ or C₂ axis) is shown. The angle of two benzene rings in KC4A-I is substantially lower (13.2°) than that of C4A (34.1°), but the other two benzene rings have a larger angle (55.7°). The distance between the centers of two opposite benzene rings in KC4A-I is 5.77 and 7.63 Å, whereas that of C4A-I is 6.83 Å. We also show the structure of H₂O•C4A complex (Fig. 4e) determined in our previous study.¹⁷ The water molecule is included inside the cup of C4A, but the tilting angles (33.2° and 37.0°) are almost the same as that of bare C4A. Bare C4A forms a hydrogen-bonded ring at the bottom of the cup with four OH groups. The OH stretching vibration of these OH groups is observed at 3158 cm⁻¹, which is substantially lower than that of phenol dimer or trimer.¹⁵ This indicates that the four OH groups are strongly hydrogen-bonded in bare C4A. The rigid cup structure of C4A is almost kept in inclusion complexes of neutral guests.¹⁵⁻¹⁸ However, the cation- π interaction between the K⁺ ion and the benzene rings highly distorts the C4A cavity to fit the cup size to the K⁺ ion and maximize the binding energy.

Figure 5 shows the comparison of the UVPD spectrum for the K⁺•C4A complex with the UV spectrum of jet-cooled C4A; these spectra are plotted as a function of UV wavenumber relative to that of the origin band.¹⁵ One can find similar spectral patterns between C4A and K⁺•C4A, namely a weak S₁-S₀ transition and strong S₂-S₀ and S₃-S₀ transitions. This similarity implies that C4A still keeps a high

symmetry upon the complex formation with K^+ , supporting the assignment of the $K^+\bullet C4A$ structure to KC4A-I (C_2 symmetry). For C4A, the S_1 state and the S_2 state, which is located $\sim 164\text{ cm}^{-1}$ above the S_1 state, have species A and E in C_4 symmetry.¹⁵ For the $K^+\bullet C4A$ complex, the origin band at 36018 cm^{-1} and the strong bands at $+138$ and $+497\text{ cm}^{-1}$ can be assigned to the electronic transitions to the $S_1(A)$, $S_2(B)$, and $S_3(B)$ states, based on the result of the TD-DFT calculation for KC4A-I. (In group theory, species E in C_4 symmetry is resolved into two species B in C_2 symmetry.) It should be noted that the S_1-S_0 transition is symmetrically allowed for both of C4A-I and KC4A-I. Very small (almost zero) S_1-S_0 oscillator strength of C4A-I is due to its conformation.¹⁵ The transition dipole moment of each phenol component (to the excited 1L_b state) is almost parallel to the benzene plane and perpendicular to its C-O bond, or perpendicular to the C_4 axis of C4A.³⁸ Therefore, the transition dipole moment of C4A with species A, which is formed by a linear combination of the transition dipole moments of the phenol parts, is very small, giving a very weak band of the $S_1(A)-S_0(A)$ transition. The UVPD spectrum of $K^+\bullet C4A$ (Fig. 5a) and the structure of KC4A-I (Fig. 4b) suggest that the above situation of the S_1-S_0 transition for C4A is likely to be kept also for $K^+\bullet C4A$. Figure 6 displays the molecular orbitals (MOs) of KC4A-I that contribute the most to each electronic transition. Similar to the case of C4A, the transition dipole moment on each phenol component seems to be almost perpendicular to the C_2 axis in $K^+\bullet C4A$; for instance, comparing the MOs of the S_2-S_0 transition (Fig. 6b) and finding the difference between them, one can recognize that the transition dipole moments of the phenol components are perpendicular to the C_2 axis. Therefore, the resulting transition dipole moment along the C_2 axis is very small, and the $S_1(A)-S_0(A)$ transition is very weak also for $K^+\bullet C4A$. As seen in Fig. 6b, the MOs localized on the pair of the benzene rings close to K^+ involve in the S_2-S_0 transition,

having a local excitation nature. In contrast, the S_3 - S_0 transition has a charge transfer character from one pair of the benzene rings to the other (Fig. 6c). Concerning the low-frequency progression with the interval of ~ 38 cm^{-1} around the origin band, a totally symmetric intermolecular vibration may involve in it. One candidate of the normal mode of KC4A-I is shown in Fig. 1S of the Supporting Information. This is the lowest-frequency vibration (47 cm^{-1}) of KC4A-I in the S_0 state and can be expressed as a large-amplitude deformation motion of the C4A cup.

4. Summary

We have constructed a mass spectrometer equipped with an ESI source and a cold, quadrupole Paul ion trap (QIT). To determine the ion temperature in our QIT, we have measured the UVPD spectra of the $\text{K}^+\bullet\text{B18C6}$ complex, and estimated the ion temperature to be ~ 10 K. We then have observed the UVPD spectrum of the $\text{K}^+\bullet\text{C4A}$ complex. The spectrum shows a very weak band at 36018 cm^{-1} and a strong one at 36156 cm^{-1} , with low-frequency (~ 38 cm^{-1}) progressions. In the geometry optimization of the $\text{K}^+\bullet\text{C4A}$ complex, we found three stable isomers. On the basis of the total energy and the results of the TD-DFT calculations, we attribute the structure of the $\text{K}^+\bullet\text{C4A}$ complex to the endo isomer (KC4A-I), in which the K^+ ion is located inside the cup. Isomer KC4A-I has C_2 symmetry, whereas bare C4A takes a C_4 structure. The symmetry reduction from C_4 to C_2 induced by the attachment of the K^+ ion will provide a different infrared spectrum other than that of bare C4A.¹⁵ Izatt and coworkers suggested that C4A derivatives show the greatest selectivity for Cs^+ among alkali metal ions.² Klinowski and coworkers concluded in their solid-state NMR studies that C4A-based molecules hold Cs^+ ion inside the cup, but Li^+ and Na^+ ions are

bonded to the phenolic oxygens.³ The application of UVPD spectroscopy to other alkali metal ion-CA complexes and the measurement of their IR-UV spectra will be of great interest and our future work.

Acknowledgment

This work is partly supported by the Japan Society for the Promotion of Science (JSPS) through the program “Strategic Young Researcher Overseas Visits Program for Accelerating Brain Circulation”. YI is grateful to Prof. Thomas R. Rizzo and Dr. Oleg V. Boyarkin in École Polytechnique Fédérale de Lausanne and Prof. Shun-ichi Ishiuchi in Tokyo Institute of Technology for their valuable comments on the ion cooling in ion traps.

Supporting Information Available: A vector drawing of the lowest-frequency vibration of KC4A-I, and a full list of authors of Ref. 35. This material is available free of charge *via* the internet at <http://pubs.acs.org>.

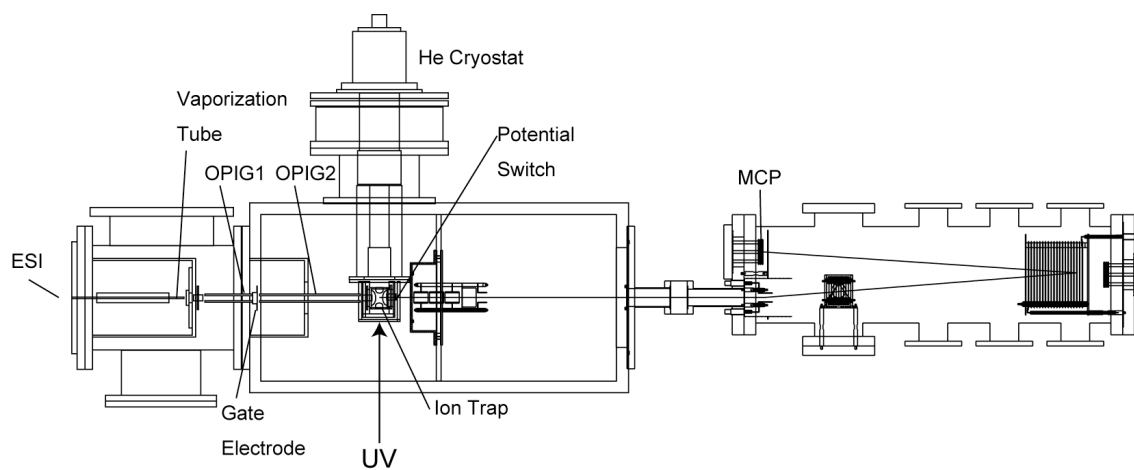


Figure 1. A schematic drawing of a mass spectrometer used in this study.

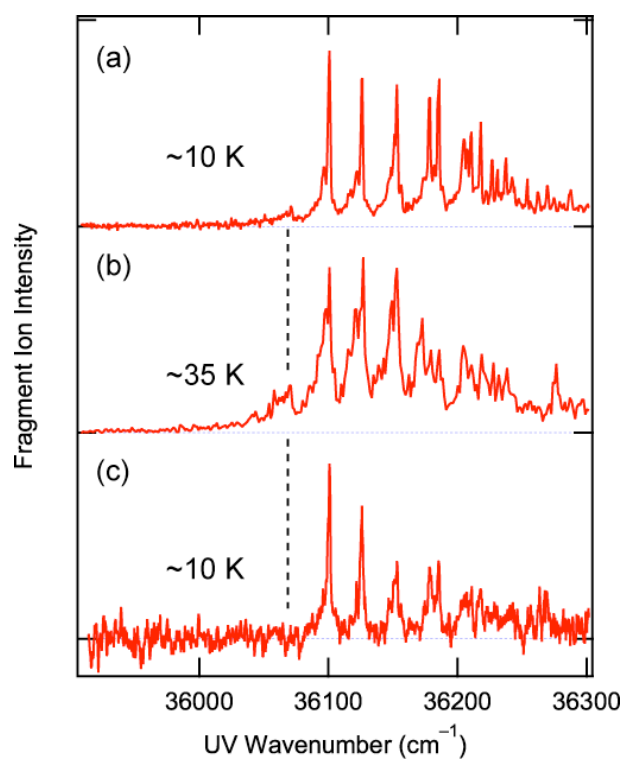


Figure 2. The UVPD spectra of the $\text{K}^+\cdot\text{B18C6}$ complex measured by using (a) a cold, 22-pole ion trap (Ref. 28) and (b, c) a cold, quadrupole Paul ion trap in this study.

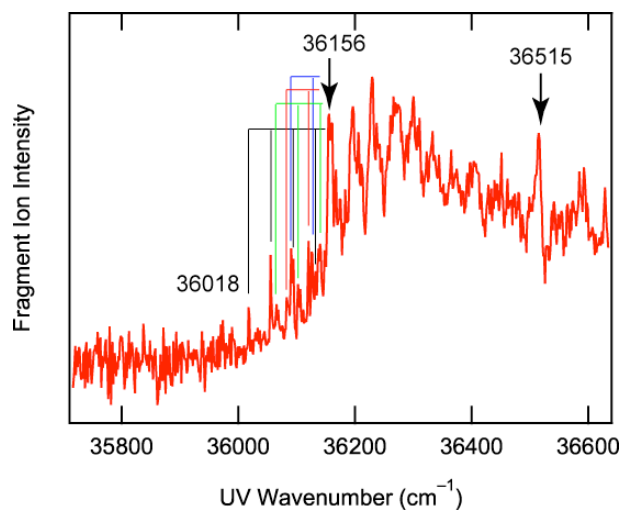


Figure 3. The UVPD spectrum of the K⁺•C4A complex. Solid lines show low-frequency progressions with an interval of ~ 38 cm⁻¹. The spectrum in the 36000–36600 cm⁻¹ region is obtained by averaging 10 repeated scans. Therefore, the fluctuation of the intensity in the region below 36000 cm⁻¹ is bigger than that above it.

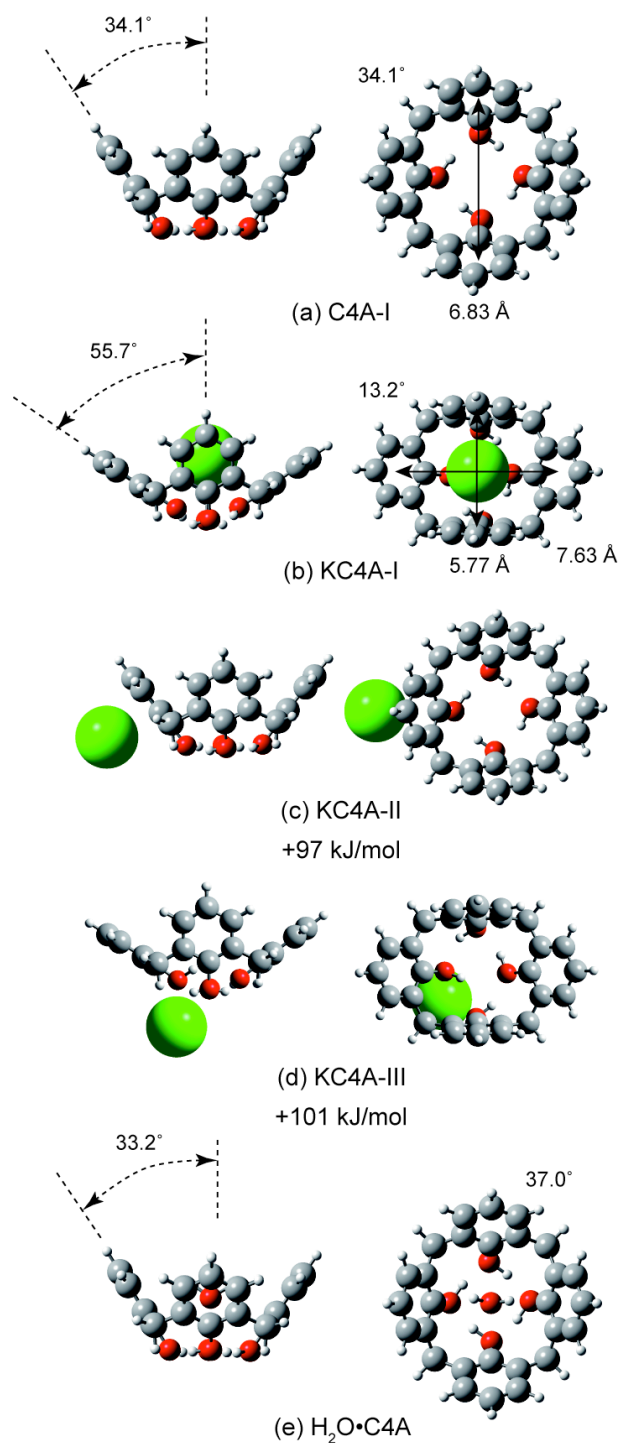


Figure 4. (a–d) The stable structures of C4A and K⁺•C4A obtained by the geometry optimization at the M05-2X/6-31+G(d) level of theory. The energy values in Figs. 4c and 4d show the total energy of isomers KC4A-II and KC4A-III relative to that of the most stable form (KC4A-I). (e) The stable structure of the H₂O•C4A complex determined in our previous study (Ref. 17).

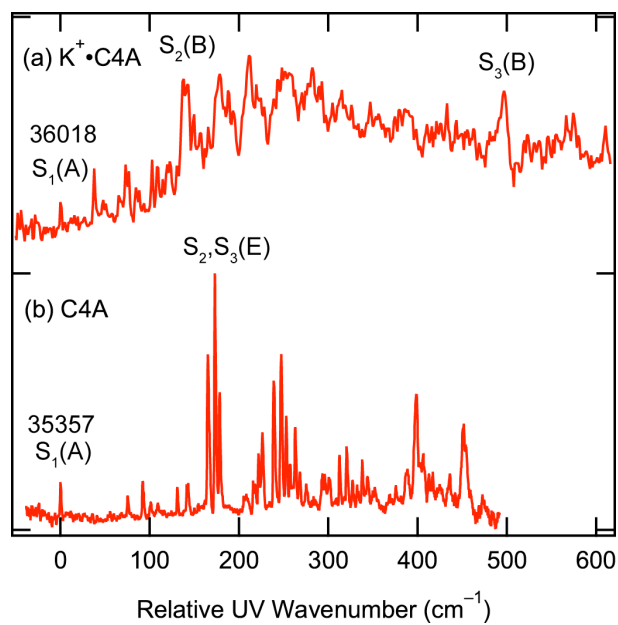


Figure 5. Comparison of the UV spectra of (a) K⁺•C4A measured in this study and (b) C4A reported in our previous study (Ref. 15).

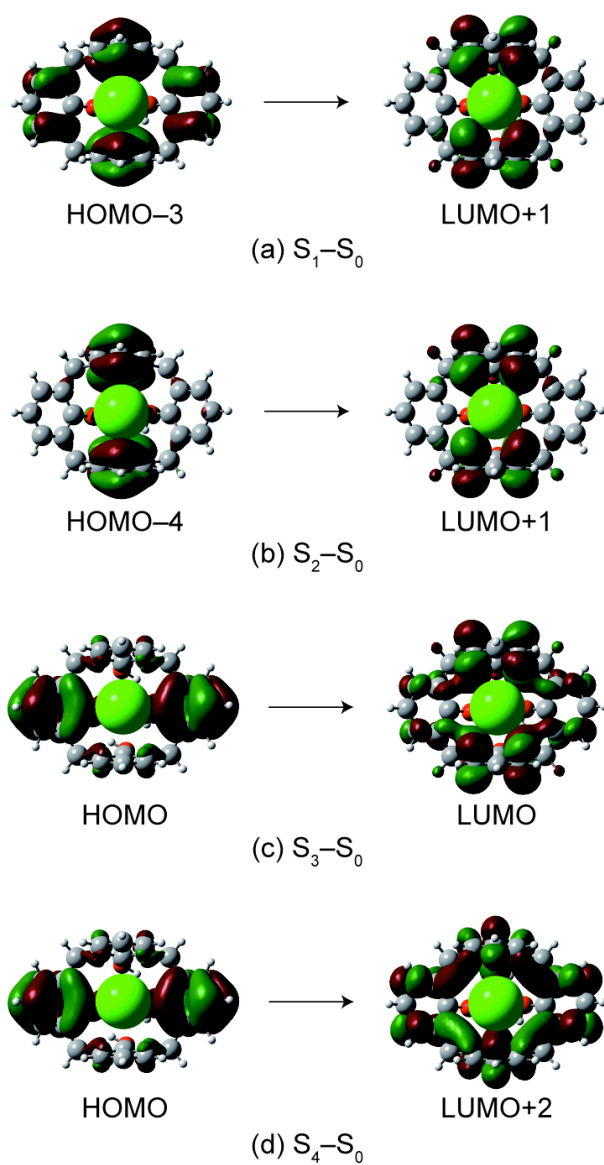


Figure 6. Molecular orbitals of isomer KC4A-I that contribute the most to each electronic transition.

Table 1. The transition energy (eV) and oscillator strength (in parentheses) from the electronic ground state for C4A and K⁺C4A calculated at the M05-2X/6-31+G(d) level of theory. The species of the electronic excited states (in parentheses, only for C4A-I and KC4A-I).

| | C4A-I ^a | KC4A-I ^a | KC4A-II | KC4A-III |
|----------------|--------------------|---------------------|---------------|---------------|
| S ₁ | 5.09 (0.0001, A) | 5.29 (0.0000, A) | 4.85 (0.0077) | 5.23 (0.0172) |
| S ₂ | 5.25 (0.1112, E) | 5.32 (0.0605, B) | 5.02 (0.0902) | 5.33 (0.0245) |
| S ₃ | 5.25 (0.1112, E) | 5.33 (0.0744, B) | 5.13 (0.0897) | 5.34 (0.0402) |
| S ₄ | 5.38 (0.0000, B) | 5.34 (0.0000, A) | 5.17 (0.0006) | 5.38 (0.0319) |

^aC4A-I and KC4A-I have C₄ and C₂ symmetry, respectively.

References

1. Gutsche, C. D., *Calixarene Revised*. Royal Society of Chemistry: Cambridge, 1998.
2. Izatt, S. R.; Hawkins, R. T.; Christensen, J. J.; Izatt, R. M., Cation Transport from Multiple Alkali Cation Mixtures Using a Liquid Membrane System Containing a Series of Calixarene Carriers. *J. Am. Chem. Soc.* **1985**, *107*, 63-66.
3. Benevelli, F.; Kolodziejcki, W.; Wozniak, K.; Klinowski, J., Solid-State Nmr Studies of Alkali Metal Ion Complexes of P-Tertbutyl-Calixarenes. *Chem. Phys. Lett.* **1999**, *308*, 65-70.
4. Schalley, C. A.; Castellano, R. K.; Brody, M. S.; Rudkevich, D. M.; Siuzdak, G.; Rebek, J., Investigating Molecular Recognition by Mass Spectrometry: Characterization of Calixarene-Based Self-Assembling Capsule Hosts with Charged Guests. *J. Am. Chem. Soc.* **1999**, *121*, 4568-4579.
5. Vincenti, M.; Irico, A., Gas-Phase Interactions of Calixarene- and Resorcinarene-Cavitands with Molecular Guests Studied by Mass Spectrometry. *Int. J. Mass spectrom.* **2002**, *214*, 23-36.
6. Weimann, D. P.; Schalley, C. A., Host-Guest Chemistry of Self-Assembling Supramolecular Capsules in the Gas Phase. *Supramol. Chem.* **2008**, *20*, 117-128.
7. Wong, P. S. H.; Yu, X.; Dearden, D. V., Complexes of P-Tert-Butylcalix[4]arene with Mono- and Dipositive Cations in the Gas Phase. *Inorg. Chim. Acta* **1996**, *246*, 259-265.
8. Zadnard, R.; Kraft, A.; Schrader, T.; Linne, U., Relative Binding Affinities of Molecular Capsules Investigated by ESI-Mass Spectrometry. *Chem. Eur. J.* **2004**, *10*, 4233-4239.
9. Ikeda, A.; Shinkai, S., On the Origin of High Ionophoricity of 1,3-Alternate Calix[4]arenes - Pi-Donor Participation in Complexation of Cations and Evidence for Metal-Tunneling through the Calix[4]arene Cavity. *J. Am. Chem. Soc.* **1994**, *116*, 3102-3110.
10. Ikeda, A.; Shinkai, S., Novel Cavity Design Using Calix[n]arene Skeletons: Toward Molecular Recognition and Metal Binding. *Chem. Rev.* **1997**, *97*, 1713-1734.
11. Iwamoto, K.; Araki, K.; Shinkai, S., Conformations and Structures of Tetra-O-Alkyl-Para-Tert-Butylcalix[4]arenes - How Is the Conformation of Calix[4]Arenes Immobilized. *J. Org. Chem.* **1991**, *56*, 4955-4962.
12. Iwamoto, K.; Shinkai, S., Syntheses and Ion Selectivity of All Conformational Isomers of Tetrakis((Ethoxycarbonyl)Methoxy)Calix[4]arene. *J. Org. Chem.* **1992**, *57*, 7066-7073.
13. Shinkai, S., Calixarenes - the 3rd-Generation of Supramolecules. *Tetrahedron* **1993**, *49*, 8933-8968.
14. Haino, T.; Hirai, E.; Fujiwara, Y.; Kashihara, K., Supramolecular Cross-Linking of [60]fullerene-Tagged Polyphenyl-Acetylene by the Host-Guest Interaction of Calix[5]arene and [60]fullerene. *Angew. Chem. Int. Ed.* **2010**, *49*, 7899-7903.
15. Ebata, T.; Hodono, Y.; Ito, T.; Inokuchi, Y., Electronic Spectra of Jet-Cooled Calix[4]arene and Its Van Der Waals Clusters: Encapsulation of a Neutral Atom in a Molecular Bowl. *J. Chem. Phys.* **2007**, *126*.
16. Ebata, T.; Hontama, N.; Inokuchi, Y.; Haino, T.; Apra, E.; Xantheas, S. S., Encapsulation of Ar_n Complexes by Calix[4]arene: Endo- Vs. Exo-Complexes. *Phys. Chem. Chem. Phys.* **2010**, *12*, 4569-4579.

17. Hontama, N.; Inokuchi, Y.; Ebata, T.; Dedonder-Lardeux, C.; Jouvet, C.; Xantheas, S. S., Structure of the Calix[4]arene-(H₂O) Cluster: The World's Smallest Cup of Water. *J. Phys. Chem. A* **2010**, *114*, 2967-2972.
18. Kaneko, S.; Inokuchi, Y.; Ebata, T.; Apra, E.; Xantheas, S. S., Laser Spectroscopic and Theoretical Studies of Encapsulation Complexes of Calix[4]arene. *J. Phys. Chem. A* **2011**, *115*, 10846-10853.
19. Boyarkin, O. V.; Mercier, S. R.; Kamariotis, A.; Rizzo, T. R., Electronic Spectroscopy of Cold, Protonated Tryptophan and Tyrosine. *J. Am. Chem. Soc.* **2006**, *128*, 2816-2817.
20. Svendsen, A.; Lorenz, U. J.; Boyarkin, O. V.; Rizzo, T. R., A New Tandem Mass Spectrometer for Photofragment Spectroscopy of Cold, Gas-Phase Molecular Ions. *Rev. Sci. Instrum.* **2010**, *81*, 073107.
21. Choi, C. M.; Baek, J. Y.; Park, K. S.; Heo, J.; Kim, N. J., Conformation-Specific Ultraviolet Spectroscopy of Benzo-18-Crown-6 Complexes with a Potassium Cation. *Chem. Phys. Lett.* **2014**, *593*, 150-153.
22. Choi, C. M.; Choi, D. H.; Heo, J.; Kim, N. J.; Kim, S. K., Ultraviolet-Ultraviolet Hole Burning Spectroscopy in a Quadrupole Ion Trap: Dibenzo[18]crown-6 Complexes with Alkali Metal Cations. *Angew. Chem. Int. Ed.* **2012**, *51*, 7297-7300.
23. Choi, C. M.; Heo, J.; Choi, M. C.; Kim, N. J., Ejection Process of Photofragment Ions from a Quadrupole Ion Trap. *Int. J. Mass spectrom.* **2013**, *337*, 12-17.
24. Choi, C. M.; Kim, H. J.; Lee, J. H.; Shin, W. J.; Yoon, T. O.; Kim, N. J.; Heo, J., Ultraviolet Photodepletion Spectroscopy of Dibenzo-18-Crown-6-Ether Complexes with Alkali Metal Cations. *J. Phys. Chem. A* **2009**, *113*, 8343-8350.
25. Choi, C. M.; Lee, J. H.; Choi, Y. H.; Kim, H. J.; Kim, N. J.; Heo, J., Ultraviolet Photodepletion Spectroscopy of Dibenzo-18-Crown-6-Ether Complexes with Alkaline Earth Metal Divalent Cations. *J. Phys. Chem. A* **2010**, *114*, 11167-11174.
26. Kim, H. J.; Shin, W. J.; Choi, C. M.; Lee, J. H.; Kim, N. J., Electronic Photodepletion Spectroscopy of Dibenzo-18-Crown-6 with a Potassium Ion. *Bull. Korean Chem. Soc.* **2008**, *29*, 1973-1976.
27. Inokuchi, Y.; Boyarkin, O. V.; Kusaka, R.; Haino, T.; Ebata, T.; Rizzo, T. R., UV and IR Spectroscopic Studies of Cold Alkali Metal Ion-Crown Ether Complexes in the Gas Phase. *J. Am. Chem. Soc.* **2011**, *133*, 12256-12263.
28. Inokuchi, Y.; Boyarkin, O. V.; Kusaka, R.; Haino, T.; Ebata, T.; Rizzo, T. R., Ion Selectivity of Crown Ethers Investigated by UV and IR Spectroscopy in a Cold Ion Trap. *J. Phys. Chem. A* **2012**, *116*, 4057-4068.
29. Inokuchi, Y.; Kusaka, R.; Ebata, T.; Boyarkin, O. V.; Rizzo, T. R., Laser Spectroscopic Study of Cold Host-Guest Complexes of Crown Ethers in the Gas Phase. *Chemphyschem* **2013**, *14*, 649-660.
30. Inokuchi, Y.; Ebata, T.; Rizzo, T. R.; Boyarkin, O. V., Microhydration Effects on the Encapsulation of Potassium Ion by Dibenzo-18-Crown-6. *J. Am. Chem. Soc.* **2014**, *136*, 1815-1824.
31. Inokuchi, Y.; Ebata, T.; Rizzo, T. R., Solvent Effects on the Encapsulation of Divalent Ions by Benzo-18-Crown-6 and Benzo-15-Crown-5. *J. Phys. Chem. A* **2015**, in press. DOI: 10.1021/acs.jpca.5b04450.
32. Kang, H.; Féraud, G.; Dedonder-Lardeux, C.; Jouvet, C., New Method for Double-Resonance Spectroscopy in a Cold Quadrupole Ion Trap and Its Application to

- UV–UV Hole-Burning Spectroscopy of Protonated Adenine Dimer. *J. Phys. Chem. Lett.* **2014**, *5*, 2760-2764.
33. Kobayashi, Y.; Inokuchi, Y.; Ebata, T., Ion Core Structure in $(\text{CS}_2)_n^+$ and $(\text{CS}_2)_n^-$ ($n = 3-10$) Studied by Infrared Photodissociation Spectroscopy. *J. Chem. Phys.* **2008**, *128*, 164319.
34. Jones, R. M.; Anderson, S. L., Simplified Radio-Frequency Generator for Driving Ion Guides, Traps, and Other Capacitive Loads. *Rev. Sci. Instrum.* **2000**, *71*, 4335-4337.
35. Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Scalmani, G.; Barone, V.; Mennucci, B.; Petersson, G. A. *et al.*, In *Gaussian 09, Revision A.1*; Gaussian, Inc.: Wallingford CT, 2009.
36. Dedonder, C.; Féraud, G.; Juvet, C., Identification of Daughter Ions through Their Electronic Spectroscopy at Low Temperature. *J. Chem. Phys.* **2014**, *141*, 131101.
37. Choi, C. M.; Choi, D. H.; Kim, N. J.; Heo, J., Effective Temperature of Protonated Tyrosine Ions in a Cold Quadrupole Ion Trap. *Int. J. Mass spectrom.* **2012**, *314*, 18-21.
38. Granucci, G.; Hynes, J. T.; Millié, P.; Tran-Thi, T.-H., A Theoretical Investigation of Excited-State Acidity of Phenol and Cyanophenols. *J. Am. Chem. Soc.* **2000**, *122*, 12243-12253.

TOC figure

